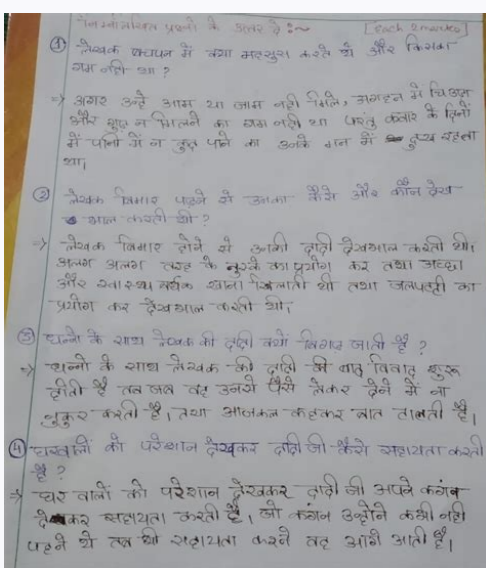
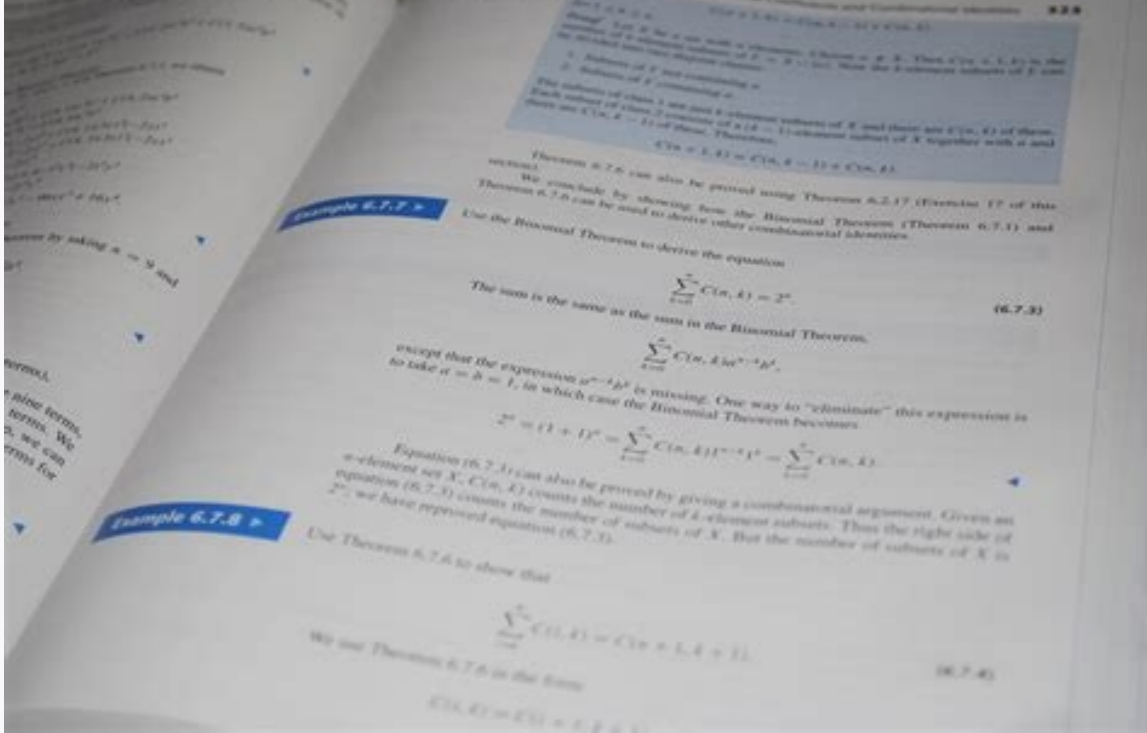


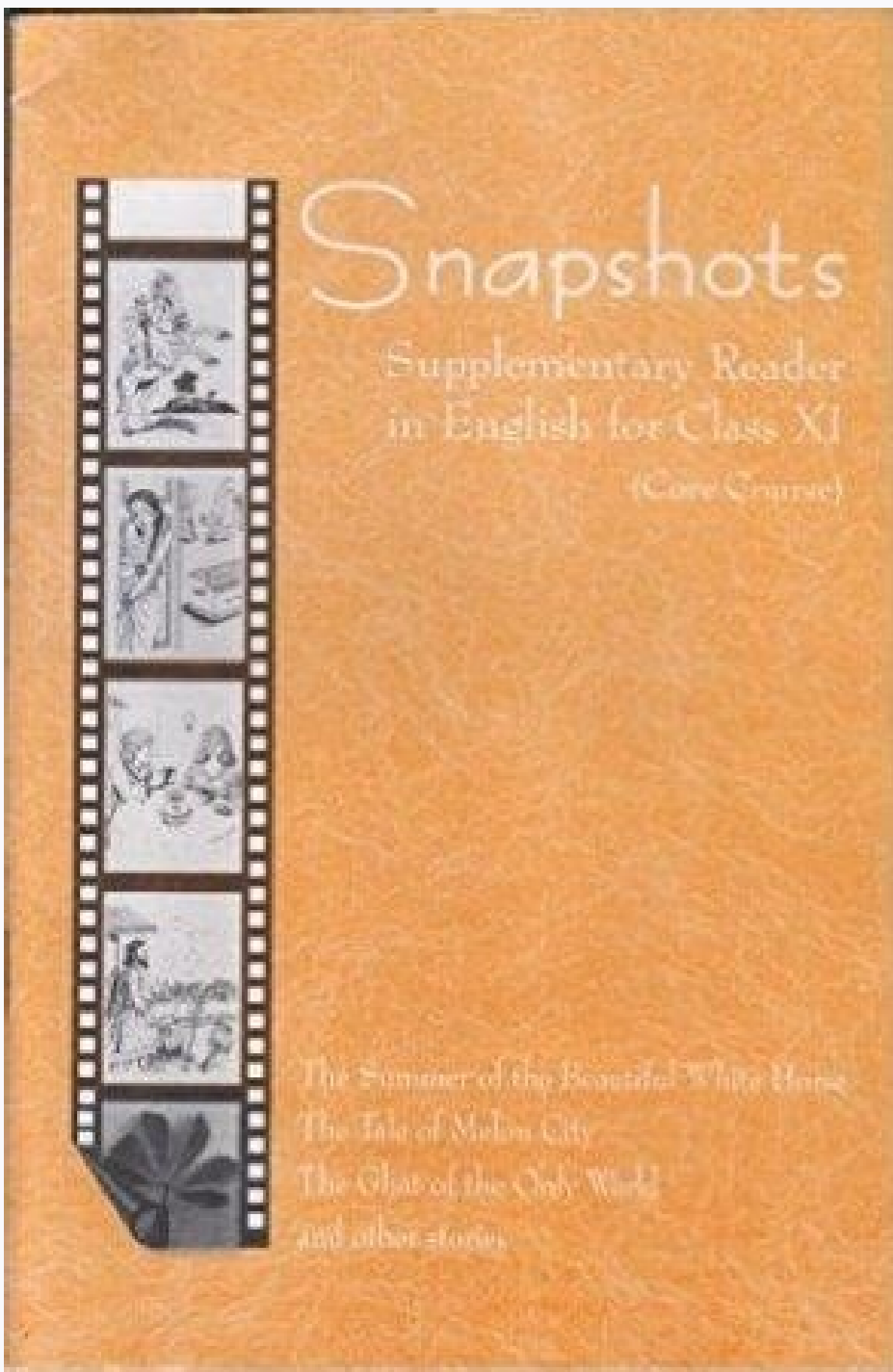
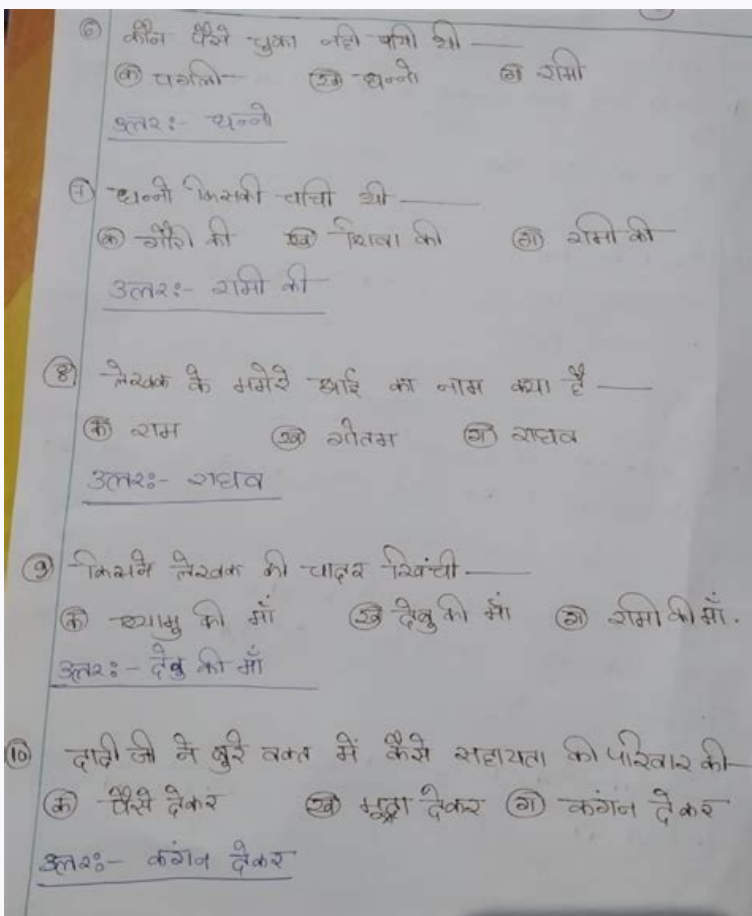
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6. a) burnt in air (1)
b) vacuum (1) solids (1)
c) Greater as a low temperature (1). Equilibrium moves to counter a lower temperature by doing the exothermic change - in this case the formation of more sulfur trioxide (1).
d) At a lower temperature, the rate of the reaction is too slow (1). 400°C is a compromise giving a good yield reasonably quickly (1).
e) Good yield even at low pressures (1). Not economically worthwhile to use higher pressures (1).
f) $\text{H}_2\text{SO}_4(l) + \text{SO}_2(g) \rightarrow \text{H}_2\text{S}_2\text{O}_7(l)$ (1)
 $\text{H}_2\text{SO}_4(l) + \text{H}_2\text{O}(l) \rightarrow 2\text{H}_2\text{SO}_4(l)$ (1)
g) See the notes - e.g. manufacture of fertilisers, manufacture of detergents, and manufacture of paper. (2 x 1)

Chemistry

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